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1) The nature of positive (canal) rays depends on
a) nature of electrode b) Nature of discharged tube
c) Nature of the residual gas d) All of mentioned

2) When 6d orbital is complete , the entering electron goes into:
a) 7f b) 6f c) 7d d) 7p

3) In the ground state of an atoms, the electron is present in:
a) nucleus b) the 1st shell c) the 2nd shell
d) nearest possible shell to the nucleus

4) The wave number of the light emitted by a certain source is $2x10^6$m. Finds its wave length
a) 500nm b) 500m c) 500µm d) $5x10^7$m.

5) Quantum number value for 2p subshell electrons are:
a) n=1, l=1 b) n=2, l=1 c) n=2, l=0 d) None

6) Orbitals having same energy are called as :
a) hybrid orbitals b) degenerate orbitals c) both d) None

7) The velocity of photon is:
a) independents of its wave length b) depends on its source
c) depend on wave length d) All

8) The negative charge on cathode rays was established by
a) Chadwick b) J. Perrin c) Goldstein d) J.J Thomason

9) The mass of neutron is :
a) a little bit smaller than that of a proton b) same is that of a proton
c) slightly higher than a proton d) All

10) The relation between energy of a photon of light and its frequency was given by
a) De-Broglie in his dual nature of matter b) Bohr in his atomic model
c) plank in his quantum theory d) All

11) Line spectra are characteristics of the
a) atoms b) molecules c) Radical d) Ions

12) The energy levels for H-atoms in Bohr models are called as
a) orbitals b) Degenerate orbitals c) orbits d) None

13) Bohr's model treated an electron only as :
a) a particles b) a wave c) both d) None

14) Energy required to ionize the hydrogen is the energy difference between two orbits and these orbits are
a) 1st and 2nd b) 1st and 3rd c) 2nd and 3rd d) 1st and infinity

15) The wave length of lyman series lies in
a) u.v region b) visible region c) I.R region d) Radio waves region

16) De Broglie's equation show the relationship of the wave length associated with the moving electron and its
a) velocity only b) Mass only c) Momentum only d) None

17) If uncertainly in position of an electron is zero, the uncertainly in its momentum would be
a) zero b) one c) $h/2\pi$ d) infinite

18) All matter particles in motion have a dual character. This statement is given by:
a) Bohr b) Dalton c) Plank d) De-Broglie

19) The frequency of a spectral line in x-ray spectrum varies as square of atomic number of element emitting it, stated as :
a) Hund b) Debroglie c) Plank d) Moseley

20) Bohr's theory cannot explain the origin of spectrum for:
a) H b) He^{+1} c) Li^{+2} d) Li

21) The credit of discovering neutron goes to the:
a) Rutherford b) Goldstein c) Chadwick d) Austen

22) The maximum number of electrons in a sub shell for which l=3 is :
a) 10 b)14 c) 6 d) 2

23) The number of electrons in the M shell of the element with atomic number 24 is
a) 24 b) 13 c) 12 d) 8

24) The symbol of the element whose atoms have the outermost electronic configuration $2S^2 2P^3$ is ,
a) N b) Li c) P d) Na

25) The number of neutron in the elements ^{9}Be is
a) 4 b) 5 c) 9 d) None

26) The electronic distribution of an atom/ion can be defined by the following
a) Aufbau principal b) Pauli's exclusion principal
c) Hund's rule d) All are used

27) Total number of possible values of —m‖ for l=2 are
a) 7 b) 5 c) 3 d) 1

28) Rutherford's experiment led to the discovery of:
a) Electron b) Proton c) Neutron d) Nucleus

29) Value of Azimuthal Quantum number of last electron in $_{11}$Na is:
a) 1 b) 2 c) 3 d) 0

30) Which of the following orbital have a dump bell shape
a) s b) d c) p d) f

31) Which series of spectral lines not lie in I.R region ?
a) Balmar series b) paschan series
c) Brackett series d) Pfund series

32) Hydrogen spectrum is an example of:
a) Atomic spectrum b) line spectrum c) Both d) none

33) Azimuthal quantum number basically tells about:
a) energy of electrons b) subshells in a shell
c) shapes of orbitals d) both b & c

34) What is formula to calculate maximum number of electrons in a subshell (s,p,d,f)?
a) (2l+1) b) 2(2l+1) c) 4(2l+2) d) All

35) Spin quantum number can only have _possible values
a) 4 b) one c) two d) zero

36) Quantum number is also called orientation quantum number:
a) principle Q.No b) Magnatic Q.No c) Spin Q. No. d) None

37) The quantum number which describes energy level, shell number and distance of electron from the nucleus is:
a) Azimuthal Q.No b) Magnatic Q.No c) Principle Q. No. d) All

38) How many quantum numbers are required to completely describe electrons?
a) one b) two c) three d) four

39) Splitting of spectral lines into constituent lines under the influence of electric field is called as:
a) Zeeman effect b) double structure c) stark effect d) All

40) Maximum No. of electrons in an orbit can be calculated by using the formula
a) (2l+1) b) $2n^2$ c) 4+2n d) All

41) Line spectrum of Na has
a) 2 white lines b) 4 green lines
c) 2 yellow lines d) 7 series of lines

42) Rainbow is an example of
a) Atomic spectrum b) line spectrum
c) continuous spectrum d) None

43) The unit of frequency is
a) waves per second b) cycles per second c) hertz d) All of mentioned

44) mvr is equal to
a) nh / π b) 2nh / π c) nh /2π d) 2nhπ / 4

45) value for planks constant (h) is
a) 6.62×10^{-30}Js b) 6.62×10^{-34} Js c) 6.02×10^{23}Js d) 9.1×10^{-31}Kgs

46) Cathode rays emitted from cathode are
(a) Canal rays (b) protons (c) electrons (d) positrons

47) Effect of magnetic field on cathode rays was studied by
(a) J-Perrin (b) J.J. Thomson (c) Stoney (d) Millikan

48) Electrons were discovered by
(a) Goldstein (b) James Chadwick (c) William Crookes (d) J.J.Thomson

49) The original glow disappears in glass discharge tube at _pressure.
(a) 0.1torr (b) 0.01torr (c)0.001torr (d) 1torr

50) 1 kg of electrons has _coulomb's of charge
(a) 1.7588×10^{11} (b) 1.7588×10^{-11} (c) 1.7588×10^{14} (d) 1.7588×10^{8}

51) The e/m value canal rays is found to be maximum for
(a) H (b)He (c) Li (d) O

52) Protons were discovered by

(a) Goldstein (b) James Chadwick (c) William Crookes (d) J.J.Thomson

53) Positive rays give flash on
(a) $AgNO_3$ plate (b)AgCl plate (c) ZnS plate (d) ZnO Plate

54) Lines of paschen series are produced when electrons jump from higher orbits to the orbit:
a) first b) second c) third d) fourth

55) Bright lines against a dark background is the characteristic of
a) Continuous spectrum b) Atomic absorption spectrum c) Atomic emission spectrum d) Line spectrum

56) Which radiations are of high frequency than others?
a) X-Rays b) UV Radiations c) Cosmic Rays d) Gamma Radiations

57) Electrons should move... nearer to the nucleus of an atom.
a) Slower b) Faster c) Constantly d) 1st slow & than fast

58) Number of waves per unit length is called as
a) Frequency b) Wave number c) Wavelength d) Photons

59) Smallest charge of electricity which has been measured so far is present on
a) Electron b) Proton c) Neutron d) Anti-neutron

60) Which rays travel in a direction opposite to cathode rays ?
a) Cathode Rays b) Canal Rays c) Neutral Rays d) Negative rays

61) Cathode Rays possess certain value of momentum & showing that these are.....
a) Radiations b) Waves c) Material Particles d) High frequency radiations

62) Second orbit of an atom is from the nucleus of an atoms than 1st.
a) 04times close b) 04times away c) 09times close d) 09times away

63) In which type of spectrum the boundary lines between the colors can not be marked?
a) Atomic spectrum b)Continuous spectrum c)Atomic absorption spectrum d)Atomic emission spectrum

64) Energy of revolving electron around a nucleus of an atom comprises...
a) K. E & P.E b) V.E & P.E c) P.E & Electrical Energy d) Kinetic & Electrical Energy

65) Gold foil used by Rutherford was of thickness
(a) 0.00004 m (b) 0.00004 cm (c) 0.00004 mm (d) 0.0000 nm

66) The tiny discrete units of light energy are called
(a) qunata (b) photons (c) wave length (d) frequency

67) Bohr explained the structure of hydrogen atom by using:
a) Rutherford's model b) Moseley's law c) Plank's quantum theory d) Discharge tubes

68) Nucleus of atom was discovered by
(a) j.j thomson (b) Rutherford (c) Max Planck's (d) Bohr

69) when alumina is placed in the path of cathode rays, then it glows
A) Yellow B) Orange C) Red D) Green

70) e/m value of cathode rays depends upon
A) Nature of gas B) Composition of glass C) Nature of cathode D) None of these

71) when a free neutron decays then it produces
A) Proton B) Electron C) Neutrino D) All of these

72) fast neutrons mostly produce
A) Alpha particles B) Beta particles C)Gamma rays D)All of these

73) The wavelength of electron moving in the first orbit of H-atom is comparable to that of
a) gamma rays b) X-rays c) cosmic rays d) radio waves

74) The wavelength of alpha particle is times smaller than that of electron.
a) 1836 b) 7344 c) 1840 d) 4473

75) Uncertainty principle is applicable to
a) Large particles b) microscopic particles c) macroscopic particles d) none of these

76) when the value of azimuthal quantum number is 1, then shape of subshell is
a) spherical b) dumbbell c) complicated d) linear

77) Higher atomic number elements give x-rays of _ wave length.
a) shorter b) greater c) discrete d) none of these

78) Bohr's model atom is contradicted by:
(A)Planck quantum theory (B)Quantization of energy of electrons (C)Heisenberg's uncertainly principle (D)Quantization of angular momentum

79) When atoms are subjected to strong electric field, spelling of spectral lines is called:
(A)Zeeman effect (B)Stark effect (C)Photoelectric effect (D)Compton effect

80) Number of bonds in N_2 molecules are:
(A) 1 (B) 2 (C) 3 (D) 4

81) An orbital which is spherical and symmetrical is:
(A)s-orbital (B)p-orbital (C)d-orbital (D)f-orbital

82) When 5d orbital is completed then entering electron goes into:
(A)6s (B)6p (C)6d (D)6f

83) What is the value of (n+l) for the 3rd sub-shell?
(A)3 (B)4 (C)5 (D)6

84) Matter is made up of extremely small particles called as:
a) Ions b) Electrons c) Atoms d) Molecules

85) Cathode rays & positive rays were discovered during:
a) Bombardment of alpha rays b) Bombardment of x-rays c) Discharge tube experiment d) Vacuum pump utility

86) Positive rays were found to contain:
a) Negatively charged particle b) Positively charged particles c) Protons & atoms d) Ions & molecules

87) Neutrons were discovered through:
a) Discharge tube experiment b) Bombardment of alpha rays on atoms of gold
b) c) Artificial radioactivity d) Rutherford's experiments

88) Which are the achievements of Rutherford's experiment?
a) Discovery of nucleus b) Presence of moving electrons around the nucleus c) Both 'a' & 'b' d) None

89) X-rays are produced when rapidly moving electrons collide with in discharge tube.
a) Heavy metal anode b) Heavy metal anti-cathode c) Both 'a' and 'b' d) Heavy metal cathode

90) Moseley discovered a simple relationship between:
a) Frequency of X-rays & mass number of target element
b) Frequency of X-rays & their energy
b) c) Frequency of X-rays & atomic number of target element
d) None

91) Wave concept of an electron was given by:
a) Davisson & Germer b) de-Broglie c) Max Planck d) Bohr

92) Concept of principal quantum number came by:
a) Heisenberg b) Bohr c) Schrodinger d) Dirac

93) The radius of 3rd orbital of hydrogen atom is
a) 13.22°A b) 2.11°A c) 4.75 °A d) 0.52°A

94) The value of vacuum permittivity _∑°' is
a) 984KJ mol^{-1} b) 63^{-8} KJ mol^{-1} c) 83.8 KJ mol^{-1} d) none of these

95) Fast neutrons have energy
a) 1.2 Mev b) 1 Mev c) 3 Mev d) 5 Mev

96) Mass of electrons in amu is
a) 1.0073 b) 1.0087 c) 5.48 $x10^{-4}$ d) 9.1 $x10^{-31}$

97) Which orbital has spherical shape
a)s b) p c) d d) f

98) If l = 2, then total number of electrons accommodated will be
a) 2 b) 6 c) 14 d) 10

99) An orbital like s, px, py, pz etc. can have at the most
a) 03electrons b). 04electrons c). 01electron d) 02electrons

100) Two electrons in the same orbital should have opposite spins is according to
a) Pauli's Exclusion principle b) Aufbau Principle c) Hund's Rule d) Moseley's law

101) The probability for finding the electron is zero between two

orbitals & this place is called as
a) Shielding effect b) Affinity place
c) Nodal plane d) Infinity point

102) For d-sub-shell there are.....of magnetic quantum number.
a) 03values b) 05values c) 01value only d) 04values

103) Electron in an orbital around the nucleus rotates about its own axis is called as
a) Clockwise rotation b) Anti-clockwise rotation
c) Self rotation d) Revolution

104) If value of l = 2 then it is called... in Azimuthal quantum number.
a) s-sub-shell b) p-sub shell c) d-sub shell d) f-sub shell

105) An electron can be described completely through its
a) 03quantum numbers b) 04quantum numbers c) 05quantum numbers d) 06quantum numbers

106) It is not possible for us to measure exact position & velocity of electron simultaneously is according to.....
a) Moseley's law b) Bohr's theory c) Plank's quantum theory
d) Heisenberg's principle

107) All material particles in motion have a dual character is according to:
a) Louis de-Broglie b) Moseley c) Lord Rutherford d) Bohr

108) Frequency of a spectral line in X-ray spectrum varies as square of atomic number of an element emitting it is
a) Heisenberg's principle b) Moseley's law
c) Hund's rule d) Pauli's exclusion principle

109) The range of visible spectrum is
(a) 300- 600nm (b) 600--- 900 nm (c) 400 ---- 750nm (d) 100 300nm

110) Spectrum obtained due to incandescent light is
(a) Continuous Spectrum (b) Line Spectrum
(c) Atomic Spectrum (d) Both b and c

111) When the electron jumps form second, third, fourth orbit to the first orbit, the transitions are known as
a) (a) Balmer series (b) Lyman series
(c) Pfund series (d) Brackett series

112) When the electron jumps form fourth, fifth and sixth orbit to the 3rd orbit, the transitions are known as
(a) Balmer series (b) Paschen series
(c) Pfund series (d) Brackett series

113) The value of Rydberg's constant is
(a) $1.60022 \times 10^{-19}\ m^{-1}$ (b) $9.1090 \times 10^{-31}\ m^{-1}$
(c) $1.0974 \times 10^{7}\ m^{-1}$ (d) $1.0974 \times 10^{-7}\ m^{-1}$

114) The limiting line of Balmer series lies in
(a) Visible region (b) UV region
(c) IR region (d) Far-IR region

115) The quantum number which describes the shape of the orbital is
a) Principle quantum number b) spin quantum number
c) Azimuthal quantum number d) magnetic quantum number

116) The value of Azimuthal number always starts from
(a) 0 (b)1 (c) 2 (d)3

117) Four d- orbitals contain four lobes while fifth contains only two lobes the orbital is
(a) d_{xy} (b) d_{xz} (c) d_{z2} (d) d_{x2-y2}

118) when tin stone is placed in the path of cathode rays, then it glows
a) A) Yellow B) Orange C) Red D) Green

119) The nature of positive rays depends upon
A) Nature of electrode B) Nature of discharge tube
C) Nature of residual gas D) Nature of voltage

120) e/m value for positive rays is found to be maximum for;
a) Hydrogen b) Oxygen c) Helium d) Nitrogen

121) The number of unpaired electrons in an atom with Z=24 at the ground state will be
a) 4 b) 3 c) 5 d) 6

122) The ionization energy of the elements of the 5th group is greater than those of the 6th group. it is because
a) 5th group element has less electronegativity
b) 5th group element have half-filled atomic orbitals
c) 6th group of the elements has a greater shielding effect
d) 5th group of the elements has a greater shielding effect

123) What is not correct for the sub-shells?
a) They are of different shapes and energies
b) They are arranged in increasing order of increasing values of (n+1)
c) They are arranged in decreasing order of energy as we move away from the nucleus
d) They are represented by azimuthal quantum numbers always

124) When a neutron decays in a nucleus of an atom
a) Mass number decreases b) atomic number increases c) Atom gets a positive charge d) All above takes place

125) Total number of electrons present in all s orbitals, all p orbitals, & d orbitals of Cs ion are respectively;
a) 6, 26, 10 b) 10, 24, 20 c) 8, 22, 24 d) 12, 20, 23

126) The total number of d-electron present in an atom of atomic number 47 is
a) 10 b) 19 c) 16 d) 20

127) Which of the following statement is wholly correct?
a) Electrons of highest energy are those nearest to the nucleus
b) All the electrons in each valence shell are defined by the same principal quantum number
c) px, py, Pz orbitals defined by the same principal quantum number differ very slightly in energy
d) There is a reasonable chance of finding an electron in the nodal region

128) How many spherical nodes are present in a 4s orbital of a hydrogen atom?
a) 0 b) 4 c) 2 d) 3

129) The number of vacant orbitals in the valence of an atom of an element with Z=14 is
a) 1 b) 3 c) 5 d) 6

130) An atom with 8 electrons in its outer shell has a tendency
a) To gain electrons b) To neither loss nor gain electrons
c) To lose electrons d) Equal to square of its amplitude

131) The total number of the fundamental particles in one atom of C-14 is
a) 6 b) 8 c) 14 d) 20

132) The d-orbital with the orientation along the x and y axes is called
a) dxy b) dy^2 c) dx^2 d) dx^2y^2

133) An atom, with atomic number 14, has a total of
a) 8 p-orbital b) 4 p-orbital c) 6 p-orbital d) 10 p-orbital

134) Which of the following orbitals are degenerate in multi-electron atoms?
a) 3d orbitals b) 1s, 2s orbitals c) 3s, 3p, 3d orbitals d) 2d orbitals

135) The arrangement of sodium, oxygen, fluorine, and strontium on the basis of increasing first ionization energy
a) Na < Sr < O < F b) Sr < Na < O < F
c) Sr < Na < F < O d) Na < Sr < F < O

136) Atomic number of an electron 'M' is 26. How many electrons are present in M-shell in ground state?
a) 11 b) 15 c) 14 d) 16

137) Angular momentum of an electron present in excited state of Hydrogen is $1.5\ h/\pi$. Electron present in its ground state?
a) 3rd orbit b) 2nd orbit c) 4th orbit d) 5th orbit

138) Which of the following electronic arrangements of the outermost orbit denotes a halogen?
a) $4s^2\ 4p^5$ b) $3s^2\ 3p^6$ c) $5s^2\ 5p^6$ d) none of these

139) Which of the following elements has the least number of electrons in its M shell?
a) Sc b) Mn c) Ni d) K

140) The very middle region in between two atomic orbitals is
a) Called nodal region and has a maximum electron density
b) Called nodal region and has a small electron density
c) Called nodal region and has 0 electron density d) Not called nodal region and has a little electron density

141) Electronic arrangement of a trivalent anion is 2, 8, 8. The atomic number of the element forming this anion is
a) 2 B) 15 c) 18 d) 19

142) What is true for the arrangement of the following sub-shells in an atom?
a) 7s>6p>5d>4f b) 4f>7s>5d>6p c) 4f>5d>6p>7s d) 5d>6p>4f>7s

143) Screening by core electrons in atoms is_?
a) More efficient than that by valence electrons b) Responsible for decrease in atomic radius going down a group
b) c) Essentially identical to that by valence electrons d) less efficient than that by valence electrons

144) Which of the following sets contains species that are isoelectronic?
a) Cl, Ar, K b) F^-, Ne, Na^+ c) Al^{3+}, S^{2-}, Ar d) P^{3-}, S^{2-}, Ar

145) The electrons in the 3rd orbit of an atom have_energy than the electrons in the 2nd orbit
a) More b) The same c) Less d) none of these

146) An atom has atomic weight of W and atomic number N then;
a) Number of electron = W– N b) Number of proton = W– N
c) Number of neutron = W– N d) Number of neutron = N

147) Which of the following element outermost orbit's last electron has magnetic quantum no. m=0?
a) Na b) O c) Cl d) N

148) All the orbitals in a given electron shell (energy level) have the same value of the _quantum number?
a) Spin b) Azimuthal c) Principle d) Magnetic

149) There are _possible values for the magnetic quantum number of an electron in a 5f subshell
a) 7 b) 14 c) 3 d) 5

150) No two electrons in same atom can have an identical set of four quantum numbers. This statement is called
a) Hund's rule b)Pauli's Exclusion principle
c) Aufbau rule d)(n+1) rule

151) Neutron is a fundamental particle that carries?
a) No charge & mass of one unit b)No charge & no mass
c)charge of +1& mass of 1 unit d)charge 1 unit & no mass

152) Principle Quantum number describes?
a) Size of orbital b) Shape of orbital
c) Orientation of an orbital d) Spin of orbital

153) An element has 18 electrons and 20 neutrons. Its charge is +2. What is its mass number?
a) 39 b) 40 c) 32 d) 38

154) Cathode rays are deflected by?
a) Both electric & magnetic field b)electric field only c)Neither electric nor magnetic field d) magnetic field only

155) Which of the following statement is wrong about electrons?
a) wave-like property b) a particle c) emits energy while moving in orbit d) motion is affected by magnetic field

156) The heaviest particle among all four given particles is?
a) Proton b) Meson c) Electron d) Neutron

157) The shape of an orbital is governed by?
a) Principle quantum number b)Spin quantum number
c)Azimuthal quantum number d)Magnetic quantum number

158) When the azimuthal quantum number l is 3 then _m' can have?
a) 5 values b) 1 value c) 7 values d) 3 value

159) Particle that most affects material properties
a) Electrons b) Protons c) Neutrons d) Valence electrons

160) The value of a magnetic quantum number of the last electron Na is?
a) 2 b) -1 c) 0 d) 3

161) A proton has approximately the same mass as
a) An electron b) An alpha particle c) A beta particle d) neutron

162) The absolute charge of an electron is _?
a) -1.6×10^{-19} C b) $+1.6 \times 10^{-19}$ C c) 1.6×10^{-19} C d) -1

163) Which one of the following sets of the quantum numbers (n, l, m, s) represents an an impossible arrangement?
a) 3, 2, -3, ½ b) 3, 2, -2, ½ c) 4, 0, 0, ½ d) 5, 3, 0, ½

164) Compared to the charge and mass of a proton, an electron has...
a) The same charge and the same mass b) An opposite charge and the same mass
c) An opposite charge and a similar mass d) The same charge and the smaller mass

165) Isotopes are:
a) Chemically similar b)Physically dissimilar c) Chemically dissimilar d) Both a and b

166) The electronic configuration of an element is $1s^2\ 2s^2\ 2p^3$. This represents a/an:
a) Ground state b)Hybridized state
c) Excited-state d) Molecular state

167) In which of the following, all have the same number of electrons:
a) Cl–, Br– and I– b)H–, H and H+
c) F–, Ne and Na^+ d) Li^+, Na^+ and K^+

168) Identify the correct order of increasing energy:
a) 1s < 2s < 3s b)2s > 3s > 1s c) 1s > 2s > 3s d) None of the above

169) The shapes of s orbitals is circular and their size:
a) Increase with the increase in principal quantum number
b)Decrease with the increase in principal quantum number
c) Remains the same with the change in principal quantum number
d) None of the above

170) The place between the two orbitals is called:
a) Free zone b)Nodal surface c) Neutral zone d) Resonance area

171) Shapes of p orbitals:
a) Circular b)Dumb-bell c) Elliptical d) Complex

172) The ion that is iso-electronic with C- atom is:
a) CN^- b)O^{+2} c) N^{+2} d) O^{-2}

173) Which of the following particles would on losing an electron has its outermost p-orbital as half-filled?
a))Nitrogen atom b)O^+ ion c) P^{-1} ion d) S^{+1} ion

174) The ionic specie having more electrons than neutrons is:
a) Mg^{+2} b)O^{2-} c) Na^+ d) F^{-1}

175) Which pair of electrons of elements will have the same chemical properties?
a) 2, 24 b)2, 4 c) 13, 22 d) 3, 11

176) An orbital which is spherically symmetrical is:
a) p-orbital b)s-orbital c) d-orbital d) f-orbital

177) The value of e/m for the electron is:
a) 1.7588×10^{11} kg C^{-1} b)1.7588×10^{-11} Ckg^{-1}
c) 1.7588×10^{11} kg C^{-1} d) 1.7588×10^{11} Ckg^{-1}

178) A nodal plane in an orbital is the plane where the electron density is:
a) Maximum b)-ve c) Zero d) +ve

179) How many times the mass of neutrons is greater than the mass of an electron?
a) 1480 b)2000 c) 1840 d) 1200

180) The mass of an oxygen atom is:
a) 2.657×10^{-23}g b)2.657×10^{23}g c) 16g d) 32g

181) The mass of a neutron is:
a) equal to proton b)A little bit smaller than proton
c) Slightly more than proton d) Slightly less than proton

182) The smallest charge of electricity which has been measured on any particle is:
a) Charge on positive ray of He b) Charge on particles
c) Charge on droplet in Millikan Experiment d) Charge on electron

183) Which of the following has highest ionization energy value?
a) Li b) Be c) H d) He

184) The mass of a proton is:
a) 9.1×10^{-10} g b)9.1×10^{-18} g c) 1.672×10^{-27} kg d) 9.1×10^{-31} kg